1. Define (in your own words):
	1. Exothermic
	2. Endothermic
	3. Activation energy
2. List 2 examples of an exothermic reaction.
3. List 2 examples of an endothermic reaction.
4. In an endothermic reaction, ∆H is \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
5. In an exothermic reaction, ∆H is \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
6. For the following examples determine whether the reaction is endothermic or exothermic.
	1. A reaction occurs in a beaker. The beaker during the reaction becomes ice cold.
	2. During a reaction the temperature of a beaker becomes extremely hot.
	3. Methane gas in a Bunsen burner produces a flame
	4. A chemical reaction proceeds and it is found that the flask has frozen to some water on the counter.
7. List and explain the factors that affect rate of reaction.